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Organic chemists use a curved arrow () to show the movement of an electron pair. A sigma(σ) bond is formed between atoms by the overlap of two atomic orbitals along the line that connects the atoms. Carbon uses sp³-hybridized orbitals to form four such bonds. These bonds are directed from the carbon nucleus toward the corners of a tetrahedron.

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Organic chemistry is a course in which one idea almost always builds on another that has gone before. It is essential, therefore, that you keep up with, or better yet, be a little ahead of your instructor. Ideally, you should try to stay one day ahead of your instructor's lectures in your own class preparations.

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In addition to ORGANIC CHEMISTRY, he is also the author or coauthor of ORGANIC CHEMISTRY: A BIOLOGICAL APPROACH, FUNDAMENTALS OF ORGANIC CHEMISTRY, THE ORGANIC CHEMISTRY OF BIOLOGICAL PATHWAYS. Customers who bought this item also bought

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Organic Chemistry Exam 1 study guide Chapter 1 Chemical bonding and chemical structure 1.2 Lewis structures Draw molecule, show all valence electrons, add electrons to give non-hydrogen atoms full octets Formal charge = (Valence electrons) - (nonbonding electrons) - (½ bonding electrons) Nonpolar when electrons are shared evenly More eneg in top right (More eneg, more attraction towards other atoms) Molecular dipole - Sum of individual dipoles 1.3 Bond length decreases left to ...

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