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**7.B Mole and Mass Avogadro's  
Number, The Mole, Grams,**

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**Atoms, Molar Mass**

**Calculations - Introduction**

Converting Between Grams and

Moles Mole Ratio Practice

Problems **Concept of Mole -**

**Part 1 | Atoms and Molecules**

**| Don't Memorise Empirical**

Formula \u0026 Molecular

Formula Determination From

Percent Composition **Molality**

**Practice Problems -**

**Molarity, Mass Percent, and**

**Density of Solution Examples**

~~Introduction to Moles~~

Stoichiometry Basic

Introduction, Mole to Mole,

Grams to Grams, Mole Ratio

Practice Problems *Mole*

*Conversions Made Easy: How*

*to Convert Between Grams and*

*Moles* How To Convert Grams

To Moles - VERY EASY!

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~~Counting Atoms: Intro to~~  
~~Moles Part 2 Converting~~  
~~Grams to Moles Using Molar~~  
~~Mass | How to Pass Chemistry~~  
~~The Mole: Avogadro's Number~~  
~~and Stoichiometry~~  
Introduction to Combustion  
Analysis, Empirical Formula  
\u0026 Molecular Formula  
Problems Dalton's Law of  
Partial Pressure Problems  
\u0026 Examples - Chemistry  
~~Stoichiometry Mole to Mole~~  
~~Conversions - Molar Ratio~~  
~~Practice Problems~~ Step by  
Step Stoichiometry Practice  
Problems | How to Pass  
Chemistry Balancing Chemical  
Equations Practice Problems  
~~Concept of Mole | Avogadro's~~  
~~Number | Atoms and Molecules~~  
~~| Don't Memorise Chemistry B~~

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CHEMISTRY WORKSHEET # 6

MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now know three things a mole can be: a molar mass,  $6.02 \times 10^{23}$  molecules and, for a gas, 22.4 liters at STP.

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CHEMISTRY B- MOLES PACKET

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CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of

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**Key** a fraction of a mole using the factor-label technique.

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Teacher Sep*

A mole of carbon contains  
 $6.02 \times 10^{23}$  atoms of carbon.

A mole of helium contains  
 $6.02 \times 10^{23}$  atoms of helium.

A mole of sodium contains  
 $6.02 \times 10^{23}$  atoms of sodium.

A mole of gold contains  $6.02$   
 $\times 10^{23}$  atoms of gold.

However, we have learned  
about seven elements that

exist as diatomic

molecules— $H_2$ ,  $N_2$ ,  $O_2$ ,  $F_2$ ,  
 $Cl_2$ ,  $Br_2$ , and  $I_2$

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4CHEMISTRY WORKSHEET # 2:

THE MOLE AS A UNIT OF

MASS Now that you know how to  
find the mass of one mole of  
a substance (molar mass) you  
can easily find the mass

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of several moles or the mass  
of a fraction of a mole  
using the factor-label  
technique. 1 mol = a molar  
mass of an atom/molecule  
(g/mol) EXAMPLE: What is the  
mass of 5.00 moles of water  
H<sub>2</sub>O? STEP 1: 2 H = 2 x (  
1.01) = 2.02 STEP 2: ? grams  
H<sub>2</sub>O = 5.00 moles H<sub>2</sub>O x  
18.02 g H<sub>2</sub>O = 90.10 ...

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6 CHEMISTRY WORKSHEET # 4  
MOLAR MASSES AND AVOGADRO'S  
NUMBER Now that you know two  
definitions of a mole (a  
molar mass and an Avogadro's  
number of particles) you

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can combine these two definitions into one problem. EXAMPLE: How many molecules are there in 90.1 grams of water?  
 $2 \text{ H} = 2 \times (1.01) = 2.02$   
 $\text{O} = 16.00$   
 $2 \text{ O} = 2 \times 16.00 = 32.00$   
 $\text{H}_2\text{O} = 2.02 + 32.00 = 34.02$   
 $90.1 \text{ g} \div 34.02 \text{ g/mol} = 2.65 \text{ mol}$   
 $2.65 \text{ mol} \times 6.02 \times 10^{23} \text{ molecules/mol} = 1.6 \times 10^{24} \text{ molecules}$   
...

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\_\_\_\_\_ PAGE 9 CHEMISTRY  
WORKSHEET # 6 MIXED MOLE  
PROBLEMS (GRAMS, MOLECULES,  
AND LITERS) You now know  
three things a mole can be:  
a molar mass,  $6.02 \times 10^{23}$   
molecules and, for a gas,  
22.4 liters at STP. We can  
use this

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CHEMISTRY WORKSHEET #5 MOLE  
PROBLEMS-MOLAR VOLUME OF A  
GAS We Have Learned That A

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Key  
Mole Is A Mass Of Material  
And Number Of Particles. A  
Mole Can Also Be A Measure  
Of Volume When We Are  
Talking About Gases. You May  
Remember From Previous  
Science Classes That All  
Gases Behave Basically The  
Same As Far As The Physical  
...

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