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Naming Acids Bases

Introduction K_a K_b

K_w pH pOH pKa pKb

H^+ OH^- Calculations -

Acids & Bases,

Buffer Solutions ,

Chemistry Review

ACIDS BASES &

SALTS-FULL

CHAPTER || CLASS

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IGCSE Chemistry:

Acids Bases and Salts

pH, pOH, H_3O^+ , OH^- ,

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~~Kw, Ka, Kb, pKa, and~~

~~pKb Basic~~

~~Calculations -Acids~~

~~and Bases Chemistry~~

~~Problems IGCSE~~

~~CHEMISTRY~~

~~REVISION [Syllabus~~

~~8] -Acids And Bases~~

~~Acid-Base Reactions~~

~~in Solution: Crash~~

~~Course Chemistry #8~~

~~Class 10th chapter 2~~

~~Acid base and salt~~

~~science PH~~

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numericals NCERT

based class UP board

2021 Exam Acids,

Bases and Salts |

Class 7 Science Sprint

for Final Exams |

Chapter 5 @Vedantu

Young Wonders HCL

- Hydrochloric Acid ||

ICSE CLASS 10

CHEMISTRY || Acids

Bases and Salts Class

10 Science Chemistry

CBSE NCERT Acids

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on Acids Bases
Introduction |
Chemistry | Don't
Memorise The
Periodic Table: Crash
Course Chemistry #4
~~Acids Bases and Salts~~
~~Le Chatelier's~~
~~Principle of Chemical~~
~~Equilibrium~~ Basic
Introduction
Calculating pH, pOH,
[H⁺], [H₃O⁺], [OH⁻] of
~~Acids and Bases~~

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~~Practice How do~~

~~Physics Wallah Earn ?~~

~~Acids and Bases, pH~~

~~and pOH How to~~

Speak Chemistrian:

Crash Course

Chemistry #11 Acids,

Bases, and pH Acids

and Bases Chemistry -

Basic Introduction

Acid-Base

Calculations Class 12

|| Chemistry || Best

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~~Aldehydes, Ketones~~

~~and Carboxylic acid~~

~~Science Notes Class~~

~~10 Chapter-2 Acids,~~

~~Bases /u0026 Salt~~

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~~Biomolecules class 12~~

~~Chemistry part 1~~

~~#NCERT Explained in~~

~~Hindi/————— Acids,~~

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Foundation - Bases

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Rao ACID BASES AND

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Acids Bases and Salts

01 : ACIDS : CBSE /

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Chem 12 Notes On

Acids

Chemistry 12 Notes

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Chem 12 Notes

on Unit 4 – Acids,

Bases and Salts

Chemistry 12 -Unit 4

-Notes Page 2 - Any

acid (weak or strong)

could have high or

low concentration.

Weak & Strong à

refers to % ionization.

Concentration à the

moles of acid

dissolved per litre.

Eg.) 10.0 M HCl à

conc. and strong [H

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Chem 12 Notes

30+] = 10.0 M 0.001

M HCl à dilute and
strong [H

D Colgur

Chem 12- Notes on
Acids & Bases - SSS
Chemistry

Acids from HClO₄ to
H₂SO₄ are 100%
ionized in water . only
solvent used in Chem
12 (and most
Chemistry) so even

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though HClO_4 is above HCl on the chart, it is no more acidic in a water solution. H_3O^+ is the strongest acid that can exist in an undissociated form in water solution. all stronger acids ionize to form H_3O^+

Chem 12- Notes on

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Chem 12 Notes

Acids & Bases - SSS

Chemistry

In this chapter, we

will focus on acids

and bases and their

chemistry. 12.1:

Introduction Certain

household chemicals,

such as some brands

of cleanser, can be

very concentrated

bases, which makes

them among the most

potentially hazardous

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Chem 12 Notes

Substances found

around the home; if
spilled on the skin,

the strong caustic
compound can

immediately remove

H^+ ions from the

flesh, resulting in

chemical burns.

12: Acids and Bases -
Chemistry LibreTexts

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Chem 12 Notes

Acids & Bases

Chemistry 12 Notes
on Unit 4 – Acids,
Bases and Salts

Chemistry 12 -Unit 4
-Notes Page 2 - Any
acid (weak or strong)
could have high or
low concentration.

Weak & Strong à
refers to % ionization.

Concentration à the
moles of acid
dissolved per litre.

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Eg.) 10.0 M HCl à
conc. and strong [H
3O+] =

Chem 12 Notes On
Acids Bases Sss

Chemistry

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Bases and Salts

Chemistry 12 -Unit 4
-Notes Page 2 - Any
acid (weak or strong)

Page 17/40

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could have high or low concentration. Weak & Strong refers to % ionization.

Concentration is the moles of acid dissolved per litre.

Eg.) 10.0 M HCl is conc. and strong $[H_3O^+] = 10.0 M$

0.001 M HCl is dilute and strong $[H_3O^+] = 0.001 M$

Notes on Acids & Bases - SSS Chemistry

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Chem 12 Notes

12.4: Acid-Base

Bases
Titrations A titration is the quantitative reaction of an acid and a base.

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Acids Bases Sss
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12 Notes On Acids
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Bases and Salts

Chemistry 12 -Unit 4

-Notes Page 2 - Any

acid (weak or strong)

could have high or

low concentration.

Weak & Strong à

refers to % ionization.

Concentration à the

moles of acid

dissolved per litre.

Eg.) 10.0 M HCl à

conc.

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Acids & Bases .

CHEMISTRY 12 UNIT

4 ACIDS, BASES AND

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Bases VIDEO. Weak
Acids VIDEO. Weak
Bases VIDEO

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know on the
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Acids and Bases 14-1
-- Acid-Base

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Bases · Bronsted-
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Bases · Proton (H⁺)
Donors and Proton
(H⁺) Acceptors

Chemistry Notes |
Chemistry Pdf --
Acids, Bases, and the
...

1. The properties of

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Chem 12 Notes

1. Acids and bases 2. pH

and pOH calculations

3. Definitions of acids
and bases 4.

Neutralization

reactions. NOTES:

Properties of acids :

- Form hydrogen ions (H^+ or protons) in solution
- Form hydronium (a water molecule bonded to a hydrogen ion shown as H_3O^+ or H^+).

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Equilibrium Unit 3

Solubility Equilibrium

Unit 4 Acids, Bases,
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Chem 12 Notes

– Acids, Bases and

Salts Notes IV.1 –

The Arrhenius Theory

Of Acids and Bases •

Acid: any substance
which releases H^+

(aq) in water. • Base:

any substance which
releases $OH^-(aq)$ in

water. • Salt: is the
neutralization

product which results
when an acid and a

base react: $HCl(aq) +$

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Chem 12 Notes

$\text{NaOH (aq) NaCl (aq) +}$
 $\text{H}_2\text{O (l) ...}$

Sss Chemistry

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Chemistry 12

Acids are classified as Organic Acids and Mineral Acids. Acids which are derived from plants and animals, they are known as Organic Acids. For Example, Citric Acid from fruit.

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Chem 12 Notes

Mineral acids are inorganic acids such as Sulphuric Acid.

They are dangerous to be used, so need more precautions.

Acids are also classified as Strong Acids or Weak Acids.

Strong acid is an acid, that completely dissociates into ions in aqueous solutions.

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Acid-Base Titrations.
Indicators. Titration
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Titration Problems
(Hebden summary)
Acid Rain. Notes from
demo: Lab
Instructions Acid-
Base Titration
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69.e), 99, 101, 104,
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PO 4) etc. Chemical

Properties of Acid: (i)

Reaction of acids with
metal: Acids give

hydrogen gas along
with respective salt

when they react with
a metal. Metal + Acid

Salt + Hydrogen

Examples: Hydrogen
gas and zinc chloride

are formed when
hydrochloric acid

reacts with zinc

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metal. Acids Bases
Sss Chemistry

Acids Bases and Salts
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I) Weak Acid

Equilibrium and K_a ;

II) Weak Base

Equilibrium and K_b ;

III) Writing Formula

(Molecular), Complete

Ionic and Net Ionic

Equations for

Acid/Base ...

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Zukowski's Class

Notes Chapter - 2

Acids, Bases and Salts

ACIDS: • These are the substances which have sour in taste. • They turns blue litmus solution to red.

• They give H^+ ions in aqueous solution

Examples of Acids : -

HCl - Hydrochloric

Acid BASES • These

substances are bitter

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in taste. • They turns red litmus solution to blue.

Chapter - 2, Acids
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