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Are Metallic Bonds?
| Properties of
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Introduction
*Hook-Ups - Types
of Chemical Bonds:*

Crash Course

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Introduction to
Ionic Bonding and
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and Metallic
Substances

**Covalent vs.
Ionic bonds CH 7
CHEMISTRY
IONIC
COMPOUNDS
(POLYATOMIC
IONS) GCSE**

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electrical

conductance in

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METALLIC

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Section 7.3 -
Bonding in Metals.
The valence
electrons of metal
atoms can be
modeled as a sea
of electrons.
Metallic bonds
consist of the
attraction of the

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free-floating
valence electrons
for the positively
charged metal
ions. Metals are
good conductors
and malleable
because of their
mobile electrons.

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Section 7.1 - Ions

OBJECTIVES:

-Determine the number of valence electrons in an atom of a representative element. 3. Section

7.1 - Ions

OBJECTIVES:

-Explain how the octet rule applies to atoms of metallic and

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nonmetallic
elements. 4.

Answer Key
*Chemistry - Chp 7 -
Ionic and Metallic
Bonding -
PowerPoint*

206 Chapter 7 •
Ionic Compounds
and Metals Section
77.1.1 Figure 7.1
As carbon dioxide
dis-solves in ocean
water, carbonate

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ions are produced.

Coral polyps capture these carbonate ions,

producing crystals of calcium

carbonate, which they secrete as an exo-skeleton. Over time, the coral reef forms. A coral reef is a complex habitat that

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Chapter 7: Ionic Compounds and Metals

Chapter 7 Ionic and Metallic Bonding

Multiple Choice

Identify the choice
that best

completes the
statement or

answers the

question. ____ 1.

What are the

valence electrons

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of an atom? a. The electrons that an atom gains when it becomes an anion
b. The electrons that remain closest to the nucleus
c. The electrons in the highest occupied energy level
d. The electrons that an atom loses when it

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and Metallic
Bonding. valence
electrons. electron
dot structures.
octet rule. halide
ions. electrons in
the highest
occupied energy
level of an

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element's....
diagrams that
show valence
electrons in the
atoms of an
elemen.... in
forming
compounds, atoms
tend to achieve the
electron confi....

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Bonding. BONDING
AND

INTERACTIONS. 7.1
Ions. Essential Under
standing Ions form

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when atoms gain or lose valence electrons, becoming electrically charged. Lesson Summary. Valence Electrons Valence electrons are the electrons in the outermost occupied energy level and are involved in ion

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formation. For a representative element, the group number equals the ...

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element.
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Types of chemical
bonds Lewis
structures and the
octet rule Ionic and
covalent bonding
Bond polarity and

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electronegativity

Formal charge

Chapter 8

Answer Key
Energetics of bond
formation and

dissociation Bond

formation: $A + B$

AB If the reaction is
energetically

favorable, AB is

lower in energy

than $A + B$. Heat is
released .

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Answer Key

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